Question 1   1 pts

How many moles are there in 25.197 g of $CaCO_3$?

Give your answer in moles and to three decimal places.

(Molar mass of $CaCO_3$ = 100.07 g mol^{-1})

Question 2   1 pts

How many moles of K^+ are there in 8.311 g of $K_2Cr_2O_7$?

Give your answer in moles and to three decimal places.

(Molar mass of $K_2Cr_2O_7$ = 294.2 g mol^{-1})

Question 3   1 pts

Question 3   1 pts

If 0.94 mol of a substance has a mass of 52.10 g, what is its molar mass?

Give your answer in g mol^{-1} and to two decimal places.

Question 4   1 pts

What is the percentage yield (no decimals in answer) where you obtained 18.6 grams from a reaction in which the theoretical yield was 25 g?

Question 5   2 pts

You dissolve 43.26 g of sodium chloride in 500 mL of water, what is the molarity of sodium cations in the solution as moles per litre?
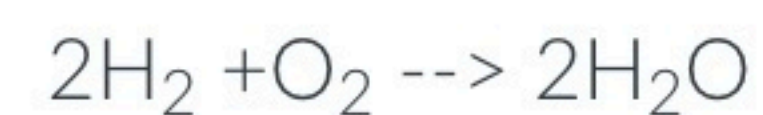
[   ]

Question 7
2 pts

As you move up the group, the [Select] ◇ number of electrons means that

the atomic radius of the atom [Select] ◇

Question 8
1 pts

For the following reaction, which would be the limiting reagent?

$2H_2 + O_2 \rightarrow 2H_2O$

You add 0.225 mol of hydrogen gas (H_2) and 0.125 mol of oxygen gas (O_2)

Question 5   2 pts

You dissolve 43.26 g of sodium chloride in 500 mL of water, what is the molarity of sodium cations in the solution as moles per litre?

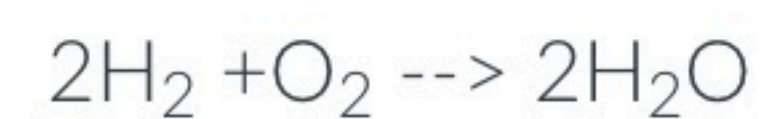## Question 6   1 pts

If you obtained a percentage yield of 39.3% from a reaction where the theoretical yield was 84 g, what was the actual amount of material you obtained in grams?

Question 8

1 pts

For the following reaction, which would be the limiting reagent?

$2H_2 + O_2 \rightarrow 2H_2O$

You add 0.225 mol of hydrogen gas (H_2) and 0.125 mol of oxygen gas (O_2)

○ Oxygen Gas (O2)

○ Hydrogen Peroxide (H2O2)

○ Hydrogen Gas (H2)

○ Water (H2O)

No new data to save. Last checked at 23:19   Submit Quiz